

9.5
10

Name: ~~XXXXXXXXXX~~
Date: _____
Period: _____

Periodic Table Practice

Use your periodic table notes to solve the following problems.

Element Close-Ups

1. Lithium

- a) Chemical Symbol: Li
- b) Atomic Number: 3
- c) Atomic Mass: 6.9
- d) Number of Protons: 3
- e) Number of Electrons: 3
- f) Number of Neutrons: 3

2. Helium

- a) Chemical Symbol: He
- b) Atomic Number: 2
- c) Atomic Mass: 4
- d) Number of Protons: 2
- e) Number of Electrons: 2
- f) Number of Neutrons: 2

3. Sulfur:

- a) Chemical Symbol: S
- b) Atomic Number: 16
- c) Atomic Mass: 32.065
- d) Number of Protons: 16
- e) Number of Electrons: 16
- f) Number of Neutrons: 16

4. Oxygen

- a) Chemical Symbol: O
- b) Atomic Number: 8
- c) Atomic Mass: 15.999
- d) Number of Protons: 8
- e) Number of Electrons: 8
- f) Number of Neutrons: 8

15

For each element, identify its period and group. Then identify the number of shells and outer electrons in the element. Finally, draw the shell diagram.

5. Boron

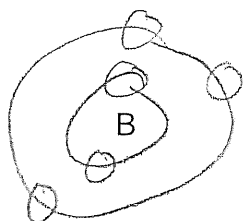
a) Period: 2

b) Group: 5

c) Number of Shells: 2

d) Number of Outer Electrons: 3

e) Draw the shell diagram



6. Magnesium

a) Period: 3

b) Group: 2

c) Number of Shells: 3

d) Number of Outer Electrons: 2

e) Draw the shell diagram

